Practice Problems in Preparation for the Quiz

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Block \_\_\_\_\_\_\_\_\_\_\_

1. For each of the following precipitation reactions, complete and balance the equation, indicating clearly which product is the precipitate. Then write the net ionic equation
2. Ba(NO3)2 + (NH4)2SO4 🡪
3. CoCl3 + NaOH 🡪
4. FeCl3 + (NH4)2S 🡪
5. For each of the following precipitation reactions, complete and balance the equation, indicating clearly which product is the precipitate. Then write the net ionic equation.
6. CaCl2 + AgC2H3O2 🡪
7. Ba(NO3)2 + NH4OH 🡪
8. NiCl2 + Na2CO3 🡪
9. Write a balance molecular equation for the precipitation reactions (if it occurs) that take place when the following aqueous solutions are mixed. Write a net ionic equation.
10. sodium sulfide and copper(II) chloride
11. potassium phosphate and aluminum chloride
12. sulfuric acid and barium chloride
13. sodium hydroxide and iron(III) chloride
14. sodium chloride and mercurous nitrate
15. potassium carbonate and chromium(III) acetate